

# Chemistry Test Cheat Sheet

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## Combustion Reaction (2 pts)

Example:  $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$

Balanced:  $C_2H_6 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

## Synthesis Reaction (2 pts)

Example:  $Na + Cl_2 \rightarrow NaCl$

Balanced:  $2Na + Cl_2 \rightarrow 2NaCl$

## Limiting/Excess Reactant (4 pts)

Example: 10g  $H_2$  & 80g  $O_2 \rightarrow$

$H_2$ : 4.95 mol,  $O_2$ : 2.50 mol

Need 5 mol  $H_2$ , have 4.95  $\rightarrow H_2$  is limiting

## Ideal Gas Law (4 pts)

$PV = nRT$

Example:  $n=1$ ,  $R=0.0821$ ,  $T=273K$ ,  $V=22.4L \rightarrow P = 1 \text{ atm}$

## % by Mass Solution (4 pts)

Formula:  $(\text{solute mass} / \text{solution mass}) \times 100$

Example: 10g NaCl in 90g water  $\rightarrow 10\%$

## EMR Problem (4 pts)

Formula:  $E = hv$ ,  $h = 6.63 \times 10^{-34}$

$v = 5.0 \times 10^{14} \text{ Hz} \rightarrow E = 3.32 \times 10^{-19} \text{ J}$

## Noble Gas Configuration (1 pt)

Sulfur (S):  $[Ne] 3s^2 3p^4$

## Formula Naming (1 pt ea.)

$Na_2SO_4$ : Sodium sulfate

$CO_2$ : Carbon dioxide

$FeCl_2$ : Iron(II) chloride

## Formula Writing (1 pt ea.)

Calcium fluoride  $\rightarrow CaF_2$

Ammonium phosphate  $\rightarrow (NH_4)_3PO_4$

Dinitrogen pentoxide  $\rightarrow N_2O_5$

## VSEPR Shape (5 pts)

Example:  $CH_4 \rightarrow$  Tetrahedral,  $109.5^\circ$

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**Jin-Ying Zhang**



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