

General Chemistry II Worksheet  
Chapters 18-19

1) Calculate the pH of a solution that is 0.322 M in nitrous acid (HNO<sub>2</sub>) and 0.178 M in potassium nitrite (KNO<sub>2</sub>).

The acid dissociation constant of nitrous acid is  $10^{-4}$   
4.50 ×

- A) 3.093                      B) 4.554                      C) 3.607                      D) 10.91                      E) 14.26

2) Which of the following acids (listed with K<sub>a</sub> values) and their conjugate base should be used to form a buffer with a pH of 2.34?

- A) C<sub>6</sub>H<sub>5</sub>OH, K<sub>a</sub> =  $1.3 \times 10^{-10}$   
B) HCN, K<sub>a</sub> =  $4.9 \times 10^{-10}$   
C) HClO<sub>2</sub>, K<sub>a</sub> =  $1.1 \times 10^{-2}$   
D) HIO, K<sub>a</sub> =  $2.3 \times 10^{-10}$   
E) HN<sub>3</sub>, K<sub>a</sub> =  $2.5 \times 10^{-5}$

3) What is the pH of the resulting solution if 45.00 mL of 0.10 M acetic acid is added to 10.00 mL of 0.10 M KOH? Assume that the volumes of the solutions are additive. K<sub>a</sub> =  $1.8 \times 10^{-5}$  for

CH<sub>3</sub>CO<sub>2</sub>H:

- A) 9.80                      B) 8.71                      C) 6.62                      D) 5.29                      E) 4.20

What is the molar solubility of silver sulfate ( Ag<sub>2</sub>SO<sub>4</sub> ) in water? The solubility-product

4) constant for Ag<sub>2</sub>SO<sub>4</sub>

is  $1.5 \times 10^{-5}$  at

- A)  $5 \times 10^{-3}$  mol/L  
B)  $5 \times 10^{-4}$  mol/L  
C)  $6 \times 10^{-2}$  mol/L  
D)  $7 \times 10^{-2}$  mol/L  
E)  $5 \times 10^{-2}$  mol/L

In which of the following solutions would solid PbCl<sub>2</sub> be expected to be least soluble at 25°C?

5) the

- A) 0.1 M NaCl                      B) 0.1 M KNO<sub>3</sub>                      C) 0.1 M HCl                      D) 0.1 M CaCl<sub>2</sub>

6) A solution contains  $2.2 \times 10^{-3}$  M in Cu<sup>2+</sup> and 0.33 M in LiCN. If the K<sub>f</sub> for Cu(CN)<sub>4</sub><sup>2-</sup> is  $1.0 \times 10^{25}$ , how much copper ion remains at equilibrium?

- A)  $1.9 \times 10^{-26}$  M  
B)  $6.7 \times 10^{-28}$  M  
C)  $2.9 \times 10^{-27}$  M  
D)  $4.6 \times 10^{-25}$  M  
E)  $3.8 \times 10^{-24}$  M

7) Which of the following processes has a  $\Delta S_{\text{sys}} > 0$ ?

- A) potassium bromide dissolves in pure water  
B)  $2 \text{NH}_3(g) + \text{CO}_2(g) \rightarrow \text{NH}_2\text{CONH}_2(aq) + \text{H}_2\text{O}(l)$   
C)  $2 \text{HBr}(g) \rightarrow \text{H}_2(g) + \text{Br}_2(l)$

# Gc2 Worksheet 6 Answers

**Christian Drosten**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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