



# 6 Hp Briggs Stratton Quantum Manual

**N Colangelo**



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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$  149.4g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}_2$  what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction  $Al + O_2 \rightarrow Al_2O_3$  4 Al 3  $O_2$  2  $Al_2O_3$  b determine the theoretical yield of  $Al_2O_3$  160.0 g Al  $\times$  1 mol Al  $\times$  2 mol  $Al_2O_3$   $\times$  101.964 g  $Al_2O_3$  302.3 g  $Al_2O_3$  26.982 g Al 4 mol Al 1 mol  $Al_2O_3$  c determine the percent yield yield actual  $\times$  100 260.0 g  $\times$  100 86.01 theoretical 302.3 g

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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