

- 13) Palladium-107, with a half-life of  $6.5 \times 10^5$  yr, undergoes  $\beta$  decay to form silver-107. How long will it take for 0.150 mol of silver-107 to form from 1.25 mol of palladium-107?  
 A)  $1.4 \times 10^7$  y  
 B)  $2.0 \times 10^7$  y  
 C)  $8.3 \times 10^5$  y  
 D)  $1.2 \times 10^5$  y
- 14) The production of nitrogen dioxide gas,  $\text{NO}_2$ , takes a three-step process.  
 $\text{NO}(\text{g}) + \text{O}(\text{g}) \rightarrow \text{NO}_2(\text{g})$   
 Multi-step process:  
 $\text{NO}(\text{g}) + \text{O}_3(\text{g}) \rightarrow \text{NO}_2(\text{g}) + \text{O}_2(\text{g}) \quad \Delta H = -198.9 \text{ kJ}$   
 $\text{O}_3(\text{g}) \rightarrow 1.5 \text{ O}_2(\text{g}) \quad \Delta H = -142.3 \text{ kJ}$   
 $\text{O}_2(\text{g}) \rightarrow 2 \text{ O}(\text{g}) \quad \Delta H = 495.0 \text{ kJ}$   
 Calculate  $\Delta H^\circ$  for the over-all reaction of nitrogen dioxide gas production from NO and O.  
 A)  $-304.1 \text{ kJ}$   
 B)  $153.8 \text{ kJ}$   
 C)  $-551.6 \text{ KJ}$   
 D)  $190.9 \text{ KJ}$
- 15) Tetraphosphorus hexaoxide,  $\text{P}_4\text{O}_6$ , is formed by the reaction of phosphorus with oxygen gas. If a mixture of 75.3 g of phosphorus and 38.7 g of oxygen produce 43.3 g of  $\text{P}_4\text{O}_6$ , what is the percent yield for the reaction?  
 A) 32.4%  
 B) 38.0%  
 C) 48.8%  
 D) 57.5%

### Analytical Chemistry

- 1) This type of Argentometric Titration uses an indirect method of determining halide ions in a sample, such that a measured excess of silver nitrate is added to the analyte solution and the excess silver is back-titrated with a thiocyanate solution.  
 A) Volhard  
 B) Mohr  
 C) Fajans  
 D) All of the above
- 2) Analysis in which the unknown is converted into a slightly soluble compound which is kept in suspension and the amount of the latter is found by comparing the intensity of the light scattered from the particles with that from a standard.  
 A) Alkalimetry  
 B) Spectrophotometry  
 C) Potentiometry  
 D) Nephelometry
- 3) Reagents that have been subjected to only superficial purification and are seldom used in analytical work.  
 A) Technical grade  
 B) Commercial grade  
 C) Analytical grade  
 D) Both A and B
- 4) An excess of a reagent is added to the sample in question and the excess reagent is titrated with another standard solution.  
 A) Direct titration  
 B) Dead stop titration  
 C) Indirect titration  
 D) Both A and B
- 5) This method is used to determine the nitrogen content of proteins and nitrogen containing organic substances.  
 A) Liebig Titration  
 B) Kjeldahl Process  
 C) Karl Fischer Titration  
 D) Saponification
- 6) The substance  $(\text{CH}_3\text{CH}_2)_2\text{NH}$  is considered  
 A) Weak acid  
 B) Weak base  
 C) Strong acid  
 D) Strong base
- 7) Which of the following is the strongest acid?  
 A)  $\text{CH}_3\text{COOH}$   
 B)  $\text{HF}$   
 C)  $\text{H}_3\text{PO}_4$   
 D)  $\text{HI}$
- 8) The acid dissociation constant  $K_a$  equals  $1.26 \times 10^{-2}$  for  $\text{HSO}_4^-$  and is  $5.6 \times 10^{-10}$  for  $\text{NH}_4^+$ . Which statement about the following equilibrium is correct?  $\text{HSO}_4^-(\text{aq}) + \text{NH}_3(\text{aq}) \leftrightarrow \text{SO}_4^{2-}(\text{aq}) + \text{NH}_4^+(\text{aq})$   
 A) The reactants will be favored because ammonia is a stronger base than the sulfate anion.  
 B) The products will be favored because the hydrogen sulfate ion is a stronger acid than the ammonium ion.  
 C) Neither reactants nor products will be favored because all of the species are weak acids or bases.  
 D) The initial concentrations of the hydrogen sulfate ion and ammonia must be known before any prediction can be made.

# Chemical Engineering Reviewer

**M Tight**



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