

Chemistry Calculations - Study-Guide

1. How many molecules are there in 2.48 grams of CO_2 ?
 B) 1.26×10^{23} molecules
2. How many formula units are there in 1.5 moles of CaCl_2 ?
 B) 9.0×10^{23}
3. How many molecules are there in 8 moles of hexagonal boron nitride, $\text{B}_3\text{N}_3\text{O}_6$?
 B) 4 molecules
4. The gram molecular mass of oxygen is 32. B) 32 g/mol
5. The gram atomic mass of oxygen is 16. B) 16 g/mol
6. What is the gram formula mass of CaCl_2 ? B) 110.98 g/mol
7. What is the number of moles in 3.6 grams of H_2O ?
 B) 2.0×10^{-1} moles of H_2O
8. The gram formula mass of CaCl_2 is 110.98 g/mol. What is the number of moles in 2.219 g of CaCl_2 ?
 C) 2.0×10^{-2} moles CaCl_2 ; 1.26×10^{22} molecules CaCl_2
 B) 2.00 g/mol
9. A 100 g sample of propanoic acid is completely combusted with 400 grams of O_2 to form a compound which is later decomposed by an unknown.
 B) 174.76 g/mol
10. What is the percent composition of aluminum in Al_2O_3 ?
 B) 52.93% Al
11. What is the empirical formula of a compound that is 80.0% carbon and 20.0% hydrogen by weight?
 B) C_4H_{10}

Chemical Quantities Study Guide Answers

James F. Hall



Chemical Quantities Study Guide Answers:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Study Guide with Selected Solutions Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

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Study Guide for Introductory Chemistry Iris Stovall, 1990

Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Denniston Katherine, 2006

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation

incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall, Donald J. DeCoste, 2000

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

A Study Guide to Chemical Principles Wilbert Hutton, 1970

Study Guide for Chemical Principles Thomas Elliott Taylor, 1979

Study Guide Darrell D. Ebbing, Susan M. Schelble, 1998

Solutions Guide, Introductory

Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry, Fourth Edition, Zumdahl James F. Hall,2000
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