

EQUILIBRIUM

- 1. The equilibrium constant, which value gives a weak description of which reaction is favored?
- 2. If a reaction is exothermic, how does the equilibrium constant change with increasing temperature?
- 3. Equilibrium is reached when the concentrations of the reactants and products are equal.
- 4. Equilibrium is reached when the concentrations of the reactants and products are equal.
- 5. Equilibrium is reached when the concentrations of the reactants and products are equal.

EMPIRICAL REPRESENTATIONS



1

For the equilibrium $A \rightleftharpoons B$, the reaction of the rate of the forward (1) & reverse (2) reaction with time is given by



MCQ TEST

EMPIRICAL REPRESENTATIONS

- 1. Equilibrium constant is constant for a reaction.
- 2. The forward reaction is favored.
- 3. The reverse reaction is favored.
- 4. The equilibrium constant is constant for a reaction.
- 5. The equilibrium constant is constant for a reaction.

2

Which of the given statements does not describe the equilibrium state properly?
 (A) The equilibrium can be approached from either direction.
 (B) The equilibrium can be imposed only if the system is an isolated system.
 (C) The free energy change of constant pressure and temperature is zero.
 (D) It is dynamic in nature.

MCQ TEST

CHEMICAL EQUILIBRIUM

- 1. Equilibrium constant is constant for a reaction.
- 2. The forward reaction is favored.
- 3. The reverse reaction is favored.
- 4. The equilibrium constant is constant for a reaction.
- 5. The equilibrium constant is constant for a reaction.

3

In the given reaction, $A + B \rightleftharpoons C$, 2 moles each of A & B are present in 10 l of volume contained in flask. It reacts at 10°C. Calculate K for the reaction.

(A) 1/3 (B) 1/4 (C) 1/5 (D) 1/6

MCQ TEST

CHEMICAL EQUILIBRIUM

EQUILIBRIUM

- 1. Equilibrium constant is constant for a reaction.
- 2. The forward reaction is favored.
- 3. The reverse reaction is favored.
- 4. The equilibrium constant is constant for a reaction.
- 5. The equilibrium constant is constant for a reaction.

4

In a given temperature, the equilibrium constant for the reaction,
 $NO_2 \rightleftharpoons \frac{1}{2} N_2O_4$ is K_1 . The equilibrium constant for the reaction,
 $2NO_2 \rightleftharpoons N_2O_4$ is K_2 . The relationship between K_1 and K_2 is
 (A) $K_2 = 4K_1$ (B) $K_2 = 2K_1$ (C) $K_2 = K_1$ (D) $K_2 = \frac{1}{4}K_1$

MCQ TEST

HOMOGENEOUS EQUILIBRIUM

- 1. Equilibrium constant is constant for a reaction.
- 2. The forward reaction is favored.
- 3. The reverse reaction is favored.
- 4. The equilibrium constant is constant for a reaction.
- 5. The equilibrium constant is constant for a reaction.

5

For the reaction,
 $Fe_3O_4 + 4CO \rightleftharpoons 3Fe + 4CO_2$
 (a) $K_1 = K_2(RT)^4$ (b) $K_1 = K_2(RT)^{-4}$
 (c) $K_1 = K_2(RT)^{-3}$ (d) $K_1 = K_2(RT)^{-1}$

MCQ TEST

LE Chatelier's Principle

- 1. Equilibrium constant is constant for a reaction.
- 2. The forward reaction is favored.
- 3. The reverse reaction is favored.
- 4. The equilibrium constant is constant for a reaction.
- 5. The equilibrium constant is constant for a reaction.

6

Which one of the following conditions will favour maximum formation of the product in the reaction $2SO_2 + O_2 \rightleftharpoons 2SO_3$, $\Delta H = -198 \text{ kJ}$
 (A) Low temperature and high pressure (B) High temperature and high pressure
 (C) Low temperature and low pressure (D) High temperature and low pressure

MCQ TEST

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