

For example:  $5P_2O_3$

P =  $\underline{(5 \times 2 =) 10}$

O =  $\underline{(5 \times 3 =) 15}$

For example:  $4Zn(NO_3)_2$

Zn =  $\underline{(4 \times 1 =) 4}$

N =  $\underline{(4 \times 1 \times 2 =) 8}$

O =  $\underline{(4 \times 3 \times 2 =) 24}$

1.  $4K_2CO_3$

K = 8

C = 4

O = 12

2.  $2Sr_3(PO_4)_2$

Sr = 6

P = 4

O = 16

3.  $3NaOH$

N =     

O =     

4.  $2(NH_4)_3N$

N = 8

H = 24

5.  $8Cl_2O$

Cl = 16

O = 8

6.  $Ca(C_2H_3O_2)_2$

Ca =     

C =     

H =     

O =     

7.  $12NaBr$

Na =     

Br =     

8.  $4Al(OH)_3$

Al =     

O =     

H =     

9.  $3NaHCO_3$

Na =     

H =     

C =     

O =     

10.  $5Ga_2(Cr_2O_7)_3$

Ga =     

Cr =     

O =     

11.  $7C_2S_2$

C =     

S =     

12.  $4Fe_2O_3$

Fe =     

O =     

13.  $6Ba(MnO_4)_2$

Ba =     

Mn =     

O =

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**Hugh Stott Taylor**



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*The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math *The Practice of Chemistry* First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes *The Practice of Chemistry* the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POC/PowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POC/PowerPoint_Final.ppt) 832KB

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**Introductory Chemistry** Darrell D. Ebbing, Albert E. Richardson, 1995

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