

Regents Chemistry Final Exam Review

Topic 4- Properties of Solutions

Define the following terms:

Saturated: Amount of solute dissolved is equal to the maximum the solvent can hold

Unsaturated: Not enough solute has been dissolved

Supersaturated: Extra solute has been dissolved

Homogeneous: The same throughout Ex: Salt water

Heterogeneous: Different throughout Ex: Oil and water

1. Explain how the following factors affect solubility:

a) Temperature

b) Pressure

c) "Like dissolves like"

2. Using Table G, answer the following questions:

a) What substances on Table G are gases?

b) How much KNO_3 is required to produce a saturated solution at $50^\circ C$ in 100 mL of water?

c) If 50 grams of HCl is added to 100 g of water at $10^\circ C$, is the solution unsaturated, saturated, or supersaturated?

d) How many grams of Sodium Nitrate is required to produce a saturated solution at $80^\circ C$ in 200 g of water?

e) KNO_3 is added to 100 mL of water to produce a saturated solution at $60^\circ C$. If the solution is cooled to $40^\circ C$, how many grams of KNO_3 will precipitate out?

Chemistry Review Topic 7 Properties Of Solutions

Ensheng Dong



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