

Chemistry 2nd Semester Final Exam Review

Chemical Bonds

1. Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.
A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.
2. Use the information from the previous question to describe the shape and volume for each phase (state) of matter.
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.
Gases particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
3. List the three phases of matter in order of increasing intermolecular attractions.
A: Gas, liquid, solid.
4. Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?
A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.

Conservation of Matter and Stoichiometry

5. The atomic mass unit and the mole are defined using what isotope? *A: Carbon-12*
6. Balance the following equations:
 - a. $Al_2O_3 + Cl_2 + C \rightarrow AlCl_3 + CO$
A: $Al_2O_3 + 3Cl_2 + 3C \rightarrow 2AlCl_3 + 3CO$
 - b. $C_4H_{10} + O_2 \rightarrow CO_2 + H_2O$
A: $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$
 - c. $P + O_2 \rightarrow P_2O_5$
A: $4P + 5O_2 \rightarrow 2P_2O_5$
7. What isotope is used to define one mole? *A: Carbon-12*
8. How many particles (atoms or molecules) are in one mole? *A: 6.02×10^{23}*
9. Find the molar mass for the following chemical formulas:
 - a. ZnF_2 *A: 103.377 g/mol*
 - b. $Al_2(SO_4)_3$ *A: 342.151 g/mol*
 - c. NH_4IO_3 *A: 192.941 g/mol*
10. How many moles are in 4.37 kg of NaOH? *A: 109.3 mol*
11. How many atoms are in 2075 g of He? *A: 3.12×10^{26} atoms*
12. What volume will 92.31 g of CO_2 occupy at STP? *A: 46.985 L*
13. How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?
 $NaOH + BaBr_2 \rightarrow NaBr + Ba(OH)_2$
A: 1.68 g NaOH
14. In the reaction, $2 Mg + O_2 \rightarrow 2 MgO$, if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?
A: 126 g MgO

Gases and Their Properties

15. What is the kinetic molecular theory (KMT)?
A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:
 - Gases are composed of a many particles that behave like hard spherical objects in a state of constant, random motion.
 - These particles move in a straight line until they collide with another particle or the walls of the container.
 - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
 - There is no force of attraction between gas particles or between the particles and the walls of the container.
 - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.

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