

# Acids and Bases

## Overview of Acids and Bases

### Acid-base Acids and Bases

Acids: Donates its proton in these processes (H<sup>+</sup>)  
 Base: Donates its lone pair in these (OH<sup>-</sup>)



Remember: It's only in water, it only HBr and HCl

What does aqueous hydrochloric acid and sodium hydroxide solution?

### Acid-base Acids and Bases

Acids are proton donors (the molecule is losing its proton)  
 Bases are proton acceptors (the molecule is still losing the proton)



### General Acids and Bases of Acid-Base Reaction

It shows the proton donor (the base) participating in the acid-base reaction and the proton acceptor (the acid) participating in the acid-base reaction.

Remember: **acid accepts lone pair** (proton donor)

Acid: Donates and Bases: Accepts (proton) Protons Protonates

### Acid-base Acids and Bases

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 Bases: Accepts its lone pair (OH<sup>-</sup>)



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## Acid Strength

### What does it mean "strong acid"?

Readily donating its proton (H<sup>+</sup>) and Bases: Accepts its lone pair (OH<sup>-</sup>)



Strong acid:  $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$   
 Weak acid:  $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$

The dissociation of the acid is a reversible reaction that is described by the equilibrium constant.



## Acid Strength and pKa

The Ka value of an acid (K<sub>a</sub>) and its conjugate base (K<sub>b</sub>) are related.



The product of the acid dissociation constant (K<sub>a</sub>) and the base dissociation constant (K<sub>b</sub>) is equal to the ion-product constant for water (K<sub>w</sub>).

$\text{pH} = -\log[\text{H}^+]$   
 $\text{pOH} = -\log[\text{OH}^-]$   
 $\text{pH} + \text{pOH} = 14$

Remember: **the stronger the acid, the lower the pKa value** (It's the opposite for the base)

Remember: **the stronger the base, the higher the pKb value**



## What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base (the base) after the acid donates its proton (H<sup>+</sup>). **The stronger the acid, the more stable its conjugate base.**



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The first factor for determining the acidity is called **inductive effect**. Inductive effect is the transmission of charge through sigma bonds.



The more electronegative the atom attached to the carbon, the stronger the acid.



The second factor for determining the acidity is called **resonance effect**. The resonance effect is the delocalization of pi electrons.



The third factor for determining the acidity is called **steric effect**. The steric effect is the hindrance of the base.

The fourth factor for determining the acidity is called **hydrogen bonding**. Hydrogen bonding is the attraction between the hydrogen of one molecule and the oxygen of another.

Remember: **the stronger the acid, the more stable its conjugate base.**



## The Position of Equilibrium for an Acid-Base Reaction

The position of equilibrium for an acid-base reaction is determined by the relative strengths of the acids and bases.



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Remember: **the stronger the acid, the more stable its conjugate base.**

## Choosing a Base to Deprotonate a Given Compound

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## Choosing an Acid to Protonate a Given Compound

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## Remember Henderson-Hasselbalch Equation

The Henderson-Hasselbalch equation relates the pH of a solution to the pKa of an acid and the ratio of the concentrations of the acid and its conjugate base.

$\text{pH} = \text{pKa} + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$

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# Chemistry Study Guide Acids And Base

**Pamela Mills,Amina El-Ashmawy**



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