Chapter 2   Diffusion in dilute solutions

1. Water evaporation

To find the rate of evaporation, we need to find the flux of water across the air film:

$$j_1 = \frac{DH}{\ell} \Delta c_1$$

Since the film is made up of air, $H = 1$. We are given D and ℓ, so we need to find Δc_1. We first calculate the difference in the partial pressure of water across the film. From a steam table we find that at 20 °C, $p_{sat} = 2.3388$ kPa. Since the air immediately above the water is saturated we have:

$$\Delta p_1 = p_{sat} - 0.5 p_{sat} = 0.5 p_{sat} = 0.5(2.3388\,kPa) = 1.1694\,kPa = 1169.4\,Pa$$

Assuming an ideal gas, we can find the concentration difference:

$$\Delta c_1 = \frac{\Delta p_1}{RT} = \frac{1169.4\,Pa}{\left(8.3145\,\frac{J}{K \cdot mol}\right)(293\,K)} = 0.48\,\frac{mol}{m^3} = 4.8 \cdot 10^{-7}\,\frac{mol}{cm^3}$$

We can now calculate the flux:

$$j_1 = \frac{0.25\,\frac{cm^2}{s}}{0.15\,cm}\left(4.8 \cdot 10^{-7}\,\frac{mol}{cm^3}\right) = 8.0 \cdot 10^{-7}\,\frac{mol}{cm^2 s}$$

To find the height change, we divide by the molar density of water:

$$\frac{8.0 \cdot 10^{-7}\,\frac{mol}{cm^2 s}}{1\,\frac{g}{cm^3} \cdot \frac{1\,mol}{18.015\,g}} = 1.44 \cdot 10^{-5}\,\frac{cm}{s} = 1.25\,\frac{cm}{day}$$

2. Diffusion across a monolayer

Recalling that the resistance is the inverse of the permeance, we have:

$$\frac{1}{P} = 2\,\frac{s}{cm} \Rightarrow P = 0.5\,\frac{cm}{s}$$

We know that the permeance is given by:

$$P = \frac{DH}{\ell} \Rightarrow D = \frac{P\ell}{H} = \frac{\left(0.5\,\frac{cm}{s}\right)(2.5 \cdot 10^{-7}\,cm)}{0.018} = 7 \cdot 10^{-6}\,\frac{cm^2}{s}$$

3. Diffusion coefficient of NO_2 in water

We treat the water as a semiinfinite slab. From Eq. (2.3-18) we know that the flux at the surface is given by:

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