

Name Key

1. Hydrogen, H₂, is used as rocket fuel. The hydrogen is burned in oxygen, O₂, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H₂ is burned?



$$\frac{200.0 \text{ g H}_2}{2.016 \text{ g H}_2} \times \frac{1 \text{ mol H}_2}{1} = 99.2 \text{ mol H}_2$$

$$99.2 \text{ mol H}_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol H}_2} = -24,006 \text{ kJ}$$

-24,006 kJ

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?

reverse so change sign of ΔH - that's in 1/2 quantity

$$\frac{1}{2} \text{CO}_2(\text{g}) + \text{SO}_2(\text{g}) \rightarrow \frac{1}{2} \text{CS}_2(\text{l}) + \frac{3}{2} \text{O}_2(\text{g})$$

$$+ \frac{1077 \text{ kJ}}{2} = 538.5 \text{ kJ}$$

3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol C}_3\text{H}_6\text{O}} \times 4.5 \text{ mol C}_3\text{H}_6\text{O} = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is 0.110 cal / g · °C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot \text{C}} \right) (225 - 25 \text{ C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N₂H₄, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



Reverse
double quantity

$$\begin{array}{r} \text{N}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \rightarrow \text{N}_2\text{H}_4(\text{l}) + \text{O}_2(\text{g}) \quad \Delta H = 622.2 \text{ kJ} \\ 2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l}) \quad \Delta H = (2)(-285.8) \text{ kJ} \\ \hline \text{N}_2(\text{g}) + 2 \text{H}_2(\text{g}) \rightarrow \text{N}_2\text{H}_4(\text{l}) \quad \Delta H = +50.6 \text{ kJ} \end{array}$$

19 3 Chemistry Practice Problems Answer Key

J Spring



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Solid-State Chemistry Ashok Kumar Jha,2023-08-04 Solid state reactions have been growing in importance in the preparation of solids crystal growth and single crystals of elements In this new volume the author discusses in depth recent advances in solid state chemistry the latest techniques of characterization of solids and several new dimensions of solid state problems Solid state reactions their kinetics phase transitions and magnetic properties are discussed in detail along with electrical properties semiconductors metals and insulators with examples and explanations Dielectric piezoelectric and ferroelectric properties of solids have been considered as well Different theories of catalysis and adsorption isotherms are covered and the author includes several experimental findings as well And of course preparation of new solids by taking advantage of recent experimental techniques have also been included in this volume This volume provides a complete description of solid state reactions electrical conductivity characterization of solids organic solids and intercalation between layers of solids The band theories of metals semiconductors and insulators are covered In characterization of solids TGA DTA DSC FTIR and PXRD techniques have been described Preparation and properties of organic solids are also explained Dielectric properties photovoltaic effects piezo electricity ferroelectricity and antiferroelectricity have been dealt with giving relatable examples A number of excitation energy levels giving rise to luminescence and photoluminescence have been explained under the topic optical processes Diamagnetic paramagnetic and ferromagnetic behaviors of solids are also explained

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Table of Contents 19 3 Chemistry Practice Problems Answer Key

1. Understanding the eBook 19 3 Chemistry Practice Problems Answer Key
 - The Rise of Digital Reading 19 3 Chemistry Practice Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Chemistry Practice Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Chemistry Practice Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Chemistry Practice Problems Answer Key
 - Personalized Recommendations
 - 19 3 Chemistry Practice Problems Answer Key User Reviews and Ratings
 - 19 3 Chemistry Practice Problems Answer Key and Bestseller Lists
5. Accessing 19 3 Chemistry Practice Problems Answer Key Free and Paid eBooks

- 19 3 Chemistry Practice Problems Answer Key Public Domain eBooks
 - 19 3 Chemistry Practice Problems Answer Key eBook Subscription Services
 - 19 3 Chemistry Practice Problems Answer Key Budget-Friendly Options
6. Navigating 19 3 Chemistry Practice Problems Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - 19 3 Chemistry Practice Problems Answer Key Compatibility with Devices
 - 19 3 Chemistry Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 19 3 Chemistry Practice Problems Answer Key
 - Highlighting and Note-Taking 19 3 Chemistry Practice Problems Answer Key
 - Interactive Elements 19 3 Chemistry Practice Problems Answer Key
8. Staying Engaged with 19 3 Chemistry Practice Problems Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Chemistry Practice Problems Answer Key
9. Balancing eBooks and Physical Books 19 3 Chemistry Practice Problems Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Chemistry Practice Problems Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Chemistry Practice Problems Answer Key
- Setting Reading Goals 19 3 Chemistry Practice Problems Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Chemistry Practice Problems Answer Key
- Fact-Checking eBook Content of 19 3 Chemistry Practice Problems Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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