Name KEY _____/20

CHEM 151 – Quiz #4
Thermo, Redox, and Electrochem
December 4, 2018

1. Consider this unbalanced reaction at 25 °C:

Zinc solid reacts with a solution containing nitrate ion, to form zinc ion and nitrogen monoxide gas.

The standard reduction potentials (E°) are:

Nitrogen half-reaction:	+ 0.957 V	red
Zinc half-reaction:	– 0.762 V	ox

a. (2 pts) Using chemical formulas, write the unbalanced reaction given above.

$$Zn_{(s)} + NO_3^-{}_{(aq)} \longrightarrow Zn^{2+}{}_{(aq)} + NO_{(g)}$$

b. (5 pts) Balance the reaction in an acidic solution using the half-reaction method practiced in class (include phases in your final balanced reaction).

$$3\left(Zn \longrightarrow Zn^{2+} + 2e^-\right) \qquad 2\left(3e^- + 4H^+ + NO_3^- \longrightarrow NO + 2H_2O\right)$$

$$3Zn + \cancel{6e^-} + 8H^+ + 2NO_3^- \longrightarrow 3Zn^{2+} + \cancel{6e^-} + 2NO + 4H_2O$$

$$3\overset{0}{Zn}_{(s)} + 8\overset{+1}{H^+}{}_{(aq)} + 2\overset{+5\ -2}{NO_3^-}{}_{(aq)} \longrightarrow 3\overset{+2}{Zn^{2+}}{}_{(aq)} + 2\overset{+2\ -2}{NO}_{(g)} + 4\overset{+1\ -2}{H_2O}_{(l)}$$

c. (3 pts) i) assign oxidation numbers to all atoms in your final balanced reaction, ii) calculate the change of oxidation state for the two atoms that underwent change, and identify if each atom underwent oxidation or reduction, and iii) label the oxidizing and reducing agents (from the reactants only).

i) see above

ii) Zn $0 \longrightarrow +2$, oxidation $\Rightarrow$ iii) $Zn_{(s)}$ = reducing agent

N $+5 \longrightarrow +2$, reduction $\Rightarrow$ $NO_3^-{}_{(aq)}$ = oxidizing agent

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