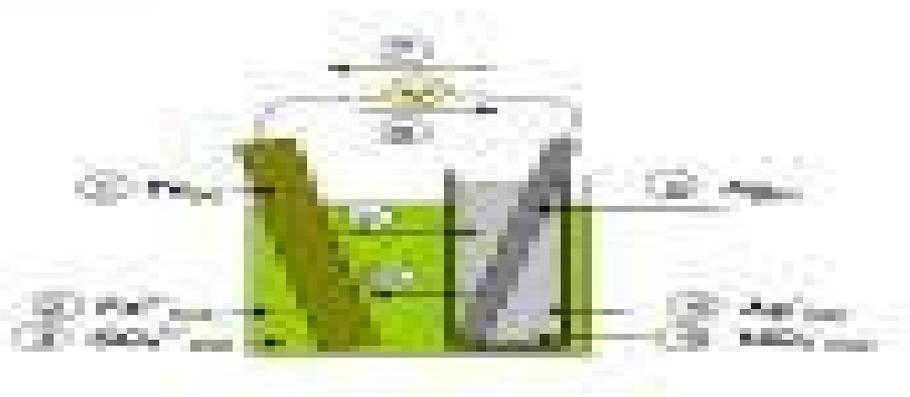


# Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. metals \_\_\_\_\_
- b. metal ions \_\_\_\_\_
- c. metals \_\_\_\_\_
- d. metal ions \_\_\_\_\_
- e. electron movement flow \_\_\_\_\_
- f. spontaneous redox reactions \_\_\_\_\_

- g. half-cell reduction half-cell \_\_\_\_\_
- h. Zn \_\_\_\_\_

i. Write overall cell reaction:  $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$  (10 marks)  
Write an equation which is spontaneous for copper & silver

\_\_\_\_\_

# Voltaic Cells Study Guide

**Mike Jess**



## Voltaic Cells Study Guide:

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