

Ch 5 Gas Law Notes

Pressure Problems:

1. How many millimeters of mercury is equal to 1.50 atmosphere?

1140 mmHg

2. Convert the pressure of 955 mm of Hg to the following units:

a. atmospheres **1.26 atm**

b. kilopascals **127 kPa**

c. pascals **127000 Pa**

3. If a force of 9.0 Newtons is applied to an area of 2.0 m^2 , what is the pressure?

4.5 Pa

4. An open ended manometer measures the atmospheric pressure to be 102 kPa. The difference in the height of the mercury is 30 mm Hg. What is the pressure exerted by the gas?

795 mmHg

5. A sample of a gas under a pressure of 822 kPa has a volume of 312 cm^3 . The pressure is increased to a 948 kPa. What is the new volume of the gas assuming that the temperature is constant? State the gas law that is applied to this problem _____.

270. cm^3 ; Boyle's Law

6. The volume of a gas is 204 cm^3 when the pressure is 925 kPa. A change in pressure caused the volume to change to 306 cm^3 . What is the pressure that causes this change in volume? State the gas law _____.

616 kPa; Boyle's Law

7. A sample of gas has a volume of 152 cm^3 when its temperature is 18°C . If the temperature is increased to 32°C , what will the new volume be assuming that the pressure is constant? State the gas law _____.

159 cm^3 ; Charles's Law

8. A sample of gas has a volume of 125 cm^3 when its temperature is 25°C . If the volume changes to 98 cm^3 , what is the new temperature? State the gas law _____.

230K; Charles's Law

9. A gas has a volume of 31.2 dm^3 at a temperature of 28°C and a pressure of 82.6 kPa. How many moles are there in the gas sample? How many molecules are in the sample? The gas is an ideal gas. State the law _____.

1.03 mol; 6.19×10^{23} molecules; ideal gas law

10. What volume will 15.0 g of carbon monoxide (CO) occupy at STP? The gas is an ideal gas. State the law _____.

12.0L; ideal gas law

Gas Law Review Answers

Terry C. Jones



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